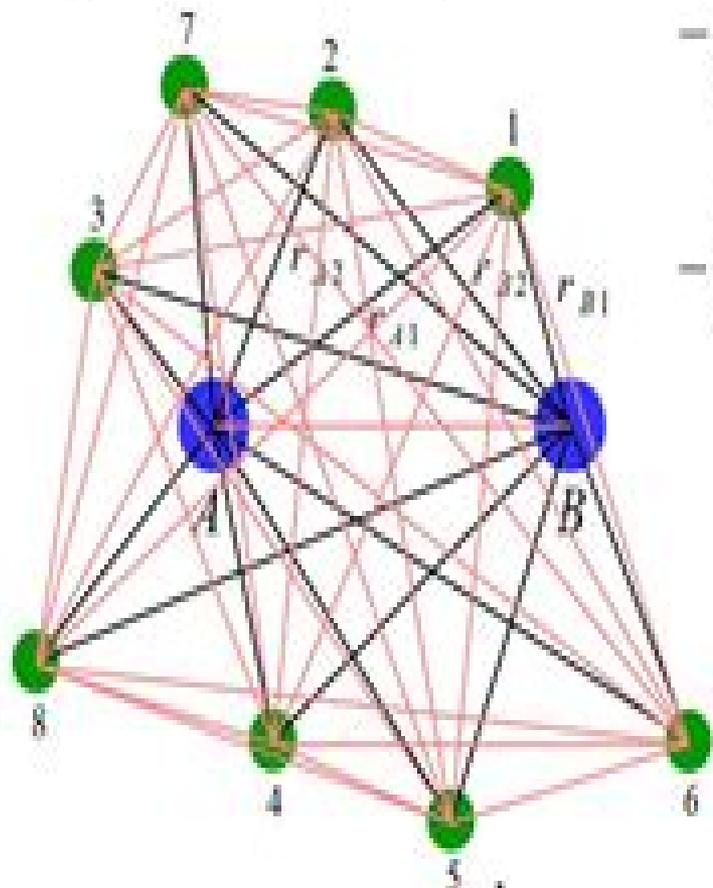


Oxygen molecule O_2



$$\frac{1}{2}N(N-1)$$

electron repulsion terms

$$-\frac{1}{2}\nabla_1^2\psi - \frac{8}{r_{A1}}\psi - \frac{8}{r_{B1}}\psi - \frac{1}{2}\nabla_2^2\psi - \frac{8}{r_{A2}}\psi - \frac{8}{r_{B2}}\psi$$

$$-\frac{1}{2}\nabla_3^2\psi - \frac{8}{r_{A3}}\psi - \frac{8}{r_{B3}}\psi - \frac{1}{2}\nabla_4^2\psi - \frac{8}{r_{A4}}\psi - \frac{8}{r_{B4}}\psi$$

and so on $+\frac{64}{R}\psi$

$$+\frac{1}{r_{12}}\psi + \frac{1}{r_{13}}\psi + \frac{1}{r_{14}}\psi + \dots + \frac{1}{r_{1,16}}\psi$$

$$+\frac{1}{r_{23}}\psi + \frac{1}{r_{24}}\psi + \frac{1}{r_{25}}\psi + \dots + \frac{1}{r_{2,16}}\psi$$

and so on

Quantum Mechanics In Chemistry

R Bogdan



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