

it to the mark, again using the boiled deionized water. Store this solution in a plastic bottle since NaOH solution dissolves glass.

Standardization of NaOH Solution

Using an analytical balance, accurately weigh 0.75 g of potassium acid phthalate (KHP) that has been dried for at least 24 hours. Record this mass reading. Place this in a 250-mL Erlenmeyer flask and add 50 mL of boiled deionized water. Add 1-2 drops of phenolphthalein indicator. Swirl to dissolve.

Titrate with the NaOH solution as your titrant until the first tinge of pink that lasts for 15 seconds is observed. Make sure to record the initial and final buret readings.

Immediately refill your buret with the titrant to minimize absorption of CO₂.

Make three trials.

$$M_{\text{NaOH}} = \frac{\text{mass}_{\text{KHP}} (1 \text{ mol NaOH})(1000)}{204.22 \frac{\text{g}}{\text{mol}} (1 \text{ mol KHP})(\text{vol NaOH})}$$

B. Analysis of the Unknown Acid

Weigh accurately 0.75 g of the provided unknown sample. Dissolve it with 50 mL boiled distilled water. Add two to three drops of phenolphthalein indicator. Titrate the solution with the standardized NaOH solution previously prepared until the end point is reached. Use the mass and volume readings to calculate for the % KHP by mass of the unknown sample.

Perform 3 trials.

$$\% \text{KHP} = \frac{\text{vol NaOH} (M \text{ NaOH})(1 \text{ mol KHP})(204.22 \frac{\text{g}}{\text{mol}})}{(1 \text{ mol NaOH})(\text{mass}_{\text{sample}})} \times 100$$

NOTE: At the end of the day, you must completely clean your buret, especially the stopcock since the basic solution, if left in, will cause it to freeze.

Post-Lab Questions: Answer briefly the questions.

1. What is titration? (2 pts)
2. In this experiment, which reagent is the titrant? The standard? The analyte? (3 pts)
3. What is standardization? When used for titration, why should a NaOH solution be standardized? (3 pts)
4. Why should cooled boiled water be used in the preparation of the NaOH solution? (2 pts)

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Table of Contents Pkg Acp Chem 128

1. Understanding the eBook Pkg Acp Chem 128
 - The Rise of Digital Reading Pkg Acp Chem 128
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Chem 128
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Chem 128
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Chem 128
 - Personalized Recommendations
 - Pkg Acp Chem 128 User Reviews and Ratings
 - Pkg Acp Chem 128 and Bestseller Lists
5. Accessing Pkg Acp Chem 128 Free and Paid eBooks
 - Pkg Acp Chem 128 Public Domain eBooks
 - Pkg Acp Chem 128 eBook Subscription Services
 - Pkg Acp Chem 128 Budget-Friendly Options
6. Navigating Pkg Acp Chem 128 eBook Formats

- ePub, PDF, MOBI, and More
 - Pkg Acp Chem 128 Compatibility with Devices
 - Pkg Acp Chem 128 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkg Acp Chem 128
 - Highlighting and Note-Taking Pkg Acp Chem 128
 - Interactive Elements Pkg Acp Chem 128
 8. Staying Engaged with Pkg Acp Chem 128
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkg Acp Chem 128
 9. Balancing eBooks and Physical Books Pkg Acp Chem 128
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Chem 128
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pkg Acp Chem 128
 - Setting Reading Goals Pkg Acp Chem 128
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pkg Acp Chem 128
 - Fact-Checking eBook Content of Pkg Acp Chem 128
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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