



AI in SaaS: Building Your Own AI-Powered SaaS Application

Proven Strategy To Create AI Powered SaaS In 2026

BATCH40 70

Minjie Lin



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

18g H₂O is your theoretical yield so percent yield = $\frac{13.7 \text{ g H}_2\text{O}}{14.9 \text{ g H}_2\text{O}} \times 100 = 91.77\%$
problems 6 & 8 are similar to problem 5

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web 1.csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation
$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
$$\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$$

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web percent yield is a measure of the efficiency of the experimental design
$$\text{Yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al₂O₃
$$\frac{160 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$$

c determine the percent yield
$$\frac{260 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$$

theoretical 302.3 g

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