

Subject:- Chemistry

# Chemical Thermodynamics and Energetics

## Practice Problems-II

Q1) Three moles of an ideal gas are expanded isothermally from a volume of 300 cm<sup>3</sup> to 2.5 L at 300 K against a pressure of 1.9 atm. Calculate the work done in L atm and joules.

Soln:-  $V_1 = 300 \text{ c.c.} = 0.3 \text{ L}$ ,  $V_2 = 2.5 \text{ L}$ ,  $T = 300 \text{ K}$ ,  $P_{ex} = 1.9 \text{ atm}$

$$W = -P_{ex}\Delta V = -P_{ex}(V_2 - V_1) \\ = -1.9(2.5 - 0.3) = -1.9 \times 2.2$$

$$\therefore W = -4.18 \text{ L atm}$$

$$\text{and } W = -4.18 \times 101.3 = -423.4 \text{ J}$$

Q2) One mole of an ideal gas is compressed from 500 cm<sup>3</sup> against a constant pressure of  $1.216 \times 10^5 \text{ Pa}$ . The work involved in the process is 36.50 J. Calculate the final volume.

Soln:-  $V_1 = 500 \text{ c.c.} = 0.5 \text{ dm}^3$ ,

$$P_{ex} = 1.216 \times 10^5 \text{ Pa} = 1.216 \times 10^2 \text{ kPa}, W = 36.50 \text{ J}$$

$$W = -P_{ex}\Delta V$$

# Problems In Chemical Thermodynamics

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