

# Stoichiometry practice problems

13. 13.2 g of iron(III) alum were dissolved in water and reduced to an iron(II) ion solution by zinc and dilute sulfuric acid. The mixture was filtered and the filtrate and washings were made up to 500 cm<sup>3</sup> in a standard volumetric flask. If 20.0 cm<sup>3</sup> of this solution required 26.5 cm<sup>3</sup> of 0.0100 mol dm<sup>-3</sup> KMnO<sub>4</sub> for oxidation:

- (a) give the ionic equation for the reduction of iron(III) ions by zinc metal.  
(b) Calculate the percentage by mass of iron in iron alum.

$$n = CV$$

$$n = \frac{m}{M_r}$$

2 : 3

Mole ratios

# Problems In Stoichiometry

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## Whispering the Strategies of Language: An Emotional Quest through **Problems In Stoichiometry**

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