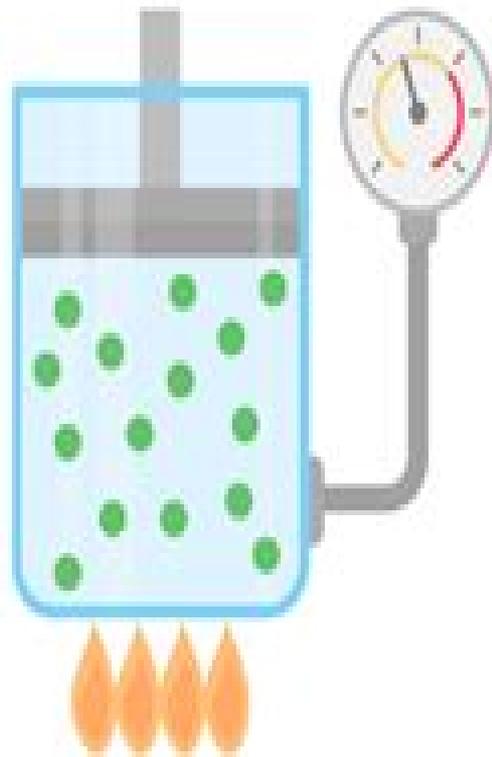


Kinetic Molecular Theory of Gases



- **Gas Particles are Small and Far Apart:** The actual volume occupied by gas molecules is negligibly small compared to the total volume of the gas.
- **Constant Random Motion:** Gas particles are in constant, random motion.
- **Elastic Collisions:** Collisions between gas particles or between particles and the container walls are perfectly elastic.
- **No Interactions:** There are no attractive or repulsive forces between the particles of an ideal gas.

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