

CHEM 131 Study Guide

Chapter 1

- ❖ Calculating density from mass and volume: $\text{Density} = \frac{\text{mass}}{\text{volume}}$
- ❖ Converting temperature from °C to K: $T \text{ (in K)} = T \text{ (in } ^\circ\text{C)} + 273.15$
- ❖ Converting temperature from K to °C: $T \text{ (in } ^\circ\text{C)} = T \text{ (in K)} - 273.15$
- ❖ Converting temperature from °C to °F: $T \text{ (in } ^\circ\text{F)} = \frac{9}{5}T \text{ (in } ^\circ\text{C)} + 32$
- ❖ Converting temperature from °F to °C: $T \text{ (in } ^\circ\text{C)} = [T \text{ (in } ^\circ\text{F)} - 32] \frac{5}{9}$

Chapter 2

- ❖ Finding the mass of an element in a given mass of compound:
Mass of element in sample = mass of compound in sample
 $\times \frac{\text{mass of element in compound}}{\text{mass of compound}}$
- ❖ Calculating the number of neutrons in an atom:
Number of neutrons = mass number - atomic number
or $N = A - Z$
- ❖ Determining the molecular mass of a formula unit of a compound:
Molecular mass = sum of atomic masses

Chapter 3

- ❖ Number of entities in one mole: 1 mol contains 6.022×10^{23} entities (to 4 sf)
- ❖ Converting amount (mol) to mass (g) using M : $\text{Mass (g)} = \text{amount (mol)} \times \frac{\text{no. of grams}}{1 \text{ mol}}$
- ❖ Converting mass (g) to amount (mol) using $1/M$: $\text{Amount (mol)} = \text{mass (g)} \times \frac{1 \text{ mol}}{\text{no. of grams}}$
- ❖ Converting amount (mol) to number of entities:
 $\text{No. of entities} = \text{amount (mol)} \times \frac{6.022 \times 10^{23} \text{ entities}}{1 \text{ mol}}$
- ❖ Converting number of entities to amount (mol):
 $\text{Amount (mol)} = \text{no. of entities} \times \frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ entities}}$
 $\text{Mass \% of element X} = \frac{\text{moles of X in formula} \times \text{molar mass of X (g/mol)}}{\text{mass (g) of 1 mol of compound}} \times 100$
- ❖ Calculating mass %:
- ❖ Finding the mass of an element in any mass of compound:
Mass of element = mass of compound
 $\times \frac{\text{mass of element in 1 mol of compound}}{\text{mass of 1 mol of compound}}$
- ❖ Calculating percent yield: $\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

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