

CHE 112 - Exam 1 Prep

• $P_i = X_i P_{\text{total}}$

• Properties of gases:

- High compressibility
- Relatively large distances between molecules
- Formation of homogenous mixtures regardless of the nature of gases

• $pV = nRT$

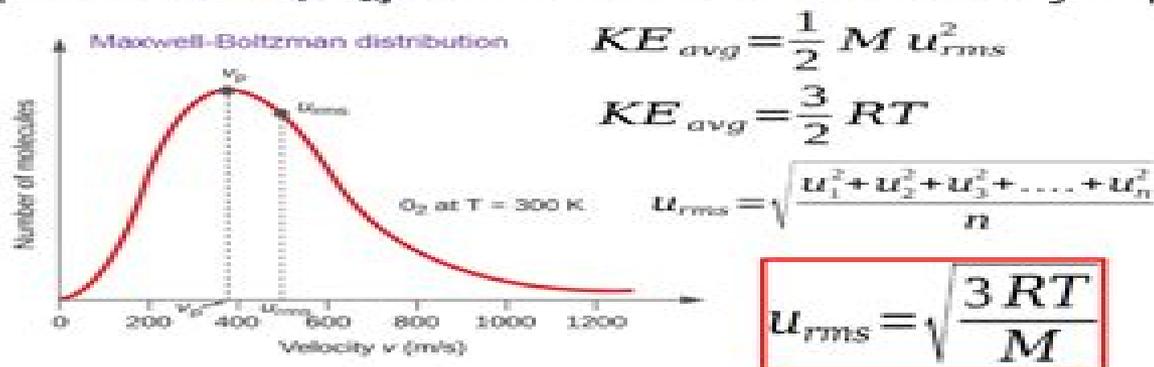
• $d = pM/RT$

• P_{total} is the sum of the pressure that each gas would exert if it were present alone

• Kinetic (molecular) Theory of Gases

- Gas molecules have mass but negligible volume
- Gas molecules are in constant motion in random direction & frequently collide with each other.
- Gas molecules exert neither attractive nor repulsive forces on one another. Collisions are perfectly elastic.
- Average Kinetic energy is determined by the temperature.

• If temperature increases, KE_{avg} increases and more molecules will have higher speeds.



• Adhesion of water to glass is stronger than cohesion within water

• Cohesive force: *intermolecular forces* present between molecules of the **same substance**; the action or property of molecules adhering to one another due to mutual attraction.

• Adhesive force: *intermolecular forces* present between **unlike molecules**. Also refers to the attractive forces between unlike substance, such as mechanical forces (sticking together) and electrostatic forces (attraction due to opposing charges). In the case of a liquid wetting agent, adhesion causes the liquid to cling to the surface on which it rests.

• Surface tension: energy required to increase the surface area by a unit amount (N/m). Surface tension is the tendency of liquid surfaces at rest to shrink into the minimum surface area possible.

• Stronger intermolecular forces = higher surface tension

• Alcohol has a lower surface tension than water

• Viscosity: the resistance of liquid to flow; depends on molar mass & type of intermolecular forces

• Ethylene glycol (HOCH₂CH₂OH) is most viscous

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