

1. (4 pts) What are the **Empirical Formula** and **Molecular Formula** for a compound that has been shown by combustion analysis to consist of 54.55% carbon, 9.09% hydrogen and 36.36% oxygen and by mass spectroscopy to have a molar mass of 132 g/mol

Empirical Formula: _____ Molecular Formula: _____

2. (4 pts) How many significant figures (sig.fig.) does each of these numbers have?

# of sig. fig.	3.1415	3.60×10^2	1200	0.003014
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3. (4 pts) Given structural formulas for 2 molecules below, write molecular and empirical formulas

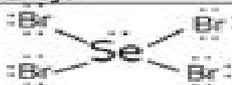
	Molecular formula	Empirical formula
		
		

4. (4 pts) Write the formula for the following ionic and molecular compounds given their name.

i) magnesium nitrate: _____ ii) dinitrogen pentoxide: _____

iii) potassium hydroxide: _____ iv) iron (III) chloride: _____

5. (4 pts) Given the following Lewis structures, **sketch** and **name** the **electron geometry** and the **molecular geometry**. For both, use wedges where appropriate to show 3-dimensionality.

			
Sketch Electron Geometry	Sketch Molecular Geometry	Sketch Electron Geometry	Sketch Molecular Geometry
Name	name	name	Name

6. (4 pts) Circle the atom or ion with the **smallest** radius from the 4 sets of data

(1) O	C	S	(2) Li	B	F
(3) Cl^-	F^-	Br^-	(4) Na^+	Mg^{2+}	K^+

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Rachel Sandford



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1. Understanding the eBook Pkg Acp Chem 111
 - The Rise of Digital Reading Pkg Acp Chem 111
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Chem 111
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Chem 111
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Chem 111
 - Personalized Recommendations
 - Pkg Acp Chem 111 User Reviews and Ratings
 - Pkg Acp Chem 111 and Bestseller Lists
5. Accessing Pkg Acp Chem 111 Free and Paid eBooks
 - Pkg Acp Chem 111 Public Domain eBooks

- Pkg Acp Chem 111 eBook Subscription Services
- Pkg Acp Chem 111 Budget-Friendly Options
- 6. Navigating Pkg Acp Chem 111 eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkg Acp Chem 111 Compatibility with Devices
 - Pkg Acp Chem 111 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkg Acp Chem 111
 - Highlighting and Note-Taking Pkg Acp Chem 111
 - Interactive Elements Pkg Acp Chem 111
- 8. Staying Engaged with Pkg Acp Chem 111
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkg Acp Chem 111
- 9. Balancing eBooks and Physical Books Pkg Acp Chem 111
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Chem 111
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkg Acp Chem 111
 - Setting Reading Goals Pkg Acp Chem 111
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkg Acp Chem 111
 - Fact-Checking eBook Content of Pkg Acp Chem 111
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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