

Experiment 17 CHE 113

1. Preparation of Saturated $\text{Ca}(\text{OH})_2$ Solution

- a. Obtain a clean beaker and fill it with approximately 100 mL of deionized (DI) water.
- b. Weigh out approximately 0.4 g of $\text{Ca}(\text{OH})_2$ and begin adding it to the water while stirring.
- c. Continue adding $\text{Ca}(\text{OH})_2$ in small increments until undissolved solid remains at the bottom, indicating a saturated solution.
- d. Allow the solution to sit and equilibrate. This solution will be used on Day 2.

1. 2. Standardization of HCl

- a. Weigh out three ~0.200 g samples of dry Na_2CO_3 and record the exact masses.
- b. Dissolve each Na_2CO_3 sample in ~20 mL of DI water in separate 100 mL beakers.
- c. Collect pH 4 and pH 10 buffer solutions to calibrate the pH probe.
- d. Set up a ring stand with a clamp, pH probe, drop counter, and burette.
- e. Set up the MeasureNet station:
 - i. MAIN MENU
 - ii. pH.mV
 - iii. pH vs. VOLUME
 - iv. CALIBRATE
 - v. Enter the room temperature
 - vi. Calibrate with pH 4 buffer. Once stable, press enter.
 - vii. Rinse the probe and calibrate with pH 10 buffer.
- f. Fill burette with 0.2 M HCl and record the initial volume.
- g. Titrate Na_2CO_3 solution while swirling. Observe equivalence point on MeasureNet.
- h. Stop MeasureNet and record the final volume. Save data as 001 and confirm with TA.
- i. Repeat steps f-h for the remaining two Na_2CO_3 solutions. Save as 002 and 003.
- j. Clean all supplies and return equipment properly.

2. 3. Determination of $[\text{OH}^-]$ in Saturated $\text{Ca}(\text{OH})_2$

- a. Filter the saturated $\text{Ca}(\text{OH})_2$ solution into a 125 mL flask using wetted filter paper.
- b. Collect fresh pH 4 and pH 10 buffers and calibrate MeasureNet again.
- c. Use the same calibration steps listed in section 2e.
- d. Measure ~25 mL of the filtered $\text{Ca}(\text{OH})_2$ solution into a 100 mL beaker and record the volume.
- e. Fill burette with 0.2 M HCl and record the initial volume.

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Didier Musso



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