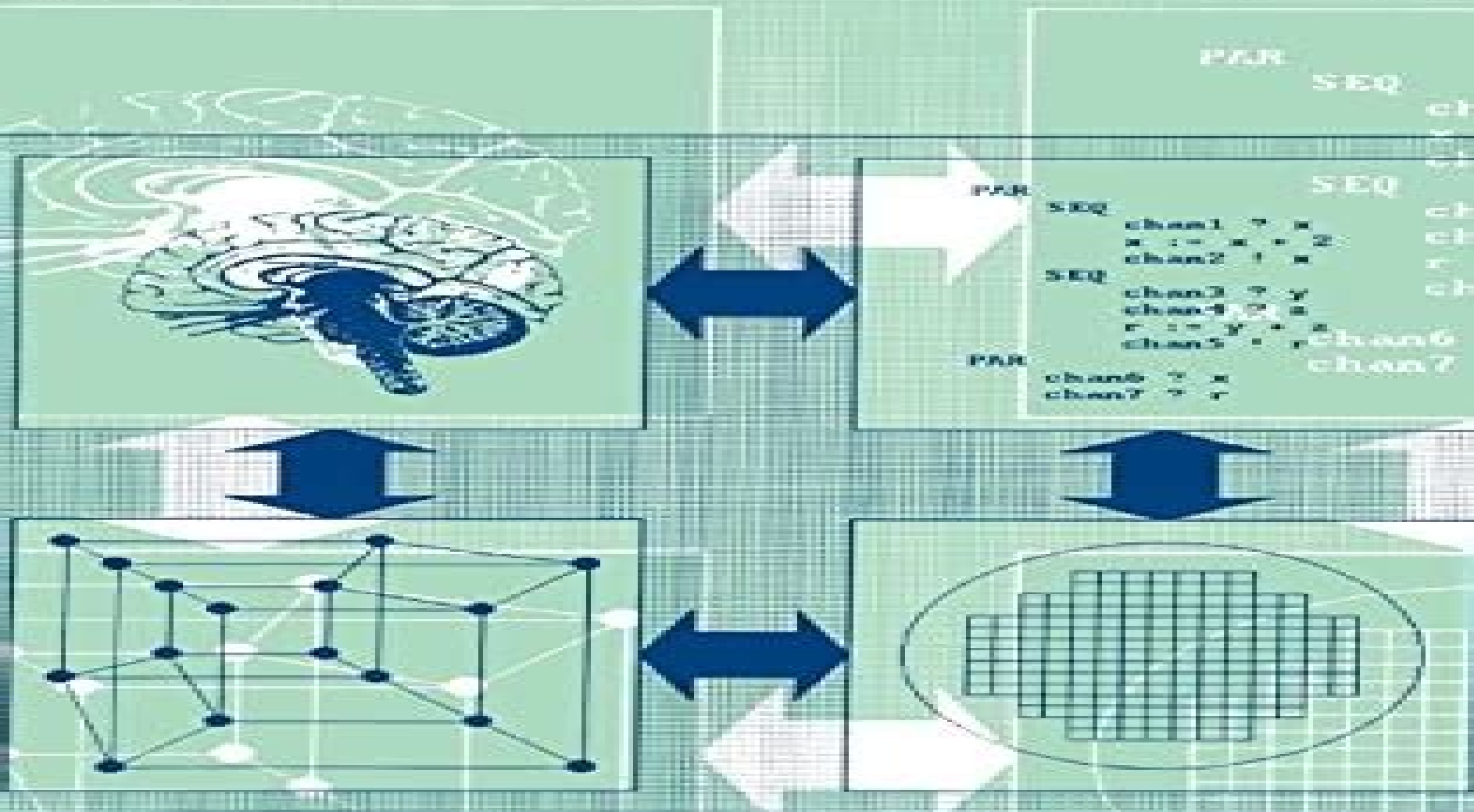


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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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molarity 0.050 M volume of filtered Ca(OH)₂ solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml OH⁻ in saturated Ca(OH)₂ M Ca in saturated Ca(OH)₂ M K_{sp} of Ca(OH)₂ average K_{sp} of Ca(OH)₂

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web the chemical equation used was Ca(OH)₂ (s) ⇌ Ca²⁺ (aq) + 2OH⁻ (aq) the solubility product was determined to be K_{sp} Ca(OH)₂ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as Ca₃(PO₄)₂ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{sp} = [Ca^{2+}]^3 [PO_4^{3-}]^2$

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