

ROSEMARY S. CAFFARELLA

A PRACTICAL
GUIDE FOR
EDUCATORS,
TRAINERS, AND
STAFF DEVELOPERS

PLANNING PROGRAMS FOR ADULT LEARNERS

SECOND EDITION

Planning Programs For Adult Learners A Practical Guide For Educators Trainers And Staff Developers

K Payea



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{ag aq} + \text{cr}_2\text{o}_7^{2-} \text{aq} \rightarrow \text{ag}_2\text{cr}_2\text{o}_7 \text{s} + 4\text{h}_2\text{o}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{na aq} + 2\text{cl aq} + \text{cu}_2\text{so}_4 \text{aq} \rightarrow 2\text{na aq} + \text{so}_4^{2-} \text{aq} + \text{cu}_2\text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{h aq} + \text{cl aq} + \text{na aq} + \text{hco}_3 \text{aq} + \text{na aq} + \text{cl aq} + \text{h}_2\text{o l} + \text{co}_2 \text{g}$ the spectator ions are cl and na net ionic equation $\text{h aq} + \text{cl aq} + \text{na aq} + \text{hco}_3 \text{aq} + \text{na aq} + \text{cl aq} + \text{h}_2\text{o l} + \text{co}_2 \text{g} \rightarrow \text{h aq} + \text{hco}_3 \text{aq} + \text{h}_2\text{o l} + \text{co}_2 \text{g}$ here is a video on spectator ions [chapter 8 04 ionic equations chemistry libretexts](#) - Mar 06 2023

web the six no_3^- ions and the six na^+ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{ba}^{2+} \text{aq} + 2\text{po}_4^{3-} \text{aq} \rightarrow \text{ba}_3\text{po}_4 \text{s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{h}^+ \text{aq} + \text{oh}^- \text{aq} \rightarrow \text{h}_2\text{o} \text{l}$ the spectator ions in $\text{koh} \text{aq}$ and $\text{hno}_3 \text{aq}$ are $\text{k}^+ \text{aq}$ and $\text{no}_3^- \text{aq}$ how to identify and list the

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web the $\text{ag}^+ \text{aq}$ and $\text{cl}^- \text{aq}$ ions become $\text{agcl} \text{s}$ but the $\text{na}^+ \text{aq}$ ions and the $\text{no}_3^- \text{aq}$ ions stay as $\text{na}^+ \text{aq}$ ions and $\text{no}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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