

QUANTITATIVE ANALYSIS

CHAPTER I INTRODUCTION*

1. Quantitative analysis is a branch of chemistry which has for its object the determination of the amount of an element, compound, or radical (usually expressed as weight per cent) that is present in a given material. It is based upon the laws of conservation of mass and of definite composition.

Various terms are used to denote the types of analysis being carried out. Thus an *ultimate* analysis is one in which the percentages of the constituent elements in the substance are found. A *proximate* analysis is the determination, at one time, of everything which reacts similarly toward a given reagent or procedure. The determination of all heavy metal impurities that are thrown down by hydrogen sulfide in a salt, or the estimation of the volatile matter in coal, is proximate analysis. Likewise an analysis may be only *partial*, such as the determination of iron in a sample of iron ore, or it may be *complete*, as in finding the exact composition of an alloy. The determination of very small percentages of substances, usually impurities, comes under the head of *trace* analysis. This must not be confused with *microanalysis*, which is the analysis of a very small sample of not over a few milligrams.

2. **Elementary quantitative analysis** is usually subdivided, for teaching purposes, into the two chief branches of *gravimetric analysis* and *volumetric analysis*. Gravimetric analysis involves the weighing of a homogeneous

* Starred chapters are used as the basis for a one-semester course.

Quantitative Analysis 2nd Edition

G Psacharopoulos



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Table of Contents Quantitative Analysis 2nd Edition

1. Understanding the eBook Quantitative Analysis 2nd Edition
 - The Rise of Digital Reading Quantitative Analysis 2nd Edition
 - Advantages of eBooks Over Traditional Books
2. Identifying Quantitative Analysis 2nd Edition
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Quantitative Analysis 2nd Edition
 - Personalized Recommendations
 - Quantitative Analysis 2nd Edition User Reviews and Ratings
 - Quantitative Analysis 2nd Edition and Bestseller Lists

5. Accessing Quantitative Analysis 2nd Edition Free and Paid eBooks
 - Quantitative Analysis 2nd Edition Public Domain eBooks
 - Quantitative Analysis 2nd Edition eBook Subscription Services
 - Quantitative Analysis 2nd Edition Budget-Friendly Options
6. Navigating Quantitative Analysis 2nd Edition eBook Formats
 - ePub, PDF, MOBI, and More
 - Quantitative Analysis 2nd Edition Compatibility with Devices
 - Quantitative Analysis 2nd Edition Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Quantitative Analysis 2nd Edition
 - Highlighting and Note-Taking Quantitative Analysis 2nd Edition
 - Interactive Elements Quantitative Analysis 2nd Edition
8. Staying Engaged with Quantitative Analysis 2nd Edition
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Quantitative Analysis 2nd Edition
9. Balancing eBooks and Physical Books Quantitative Analysis 2nd Edition
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Quantitative Analysis 2nd Edition
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Quantitative Analysis 2nd Edition
 - Setting Reading Goals Quantitative Analysis 2nd Edition
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Quantitative Analysis 2nd Edition
 - Fact-Checking eBook Content of Quantitative Analysis 2nd Edition
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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