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**Anthony S. Fauci**



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absorbed or released in the reaction what is  $\Delta H$  for the reverse reaction what is  $\Delta H$  when 5.35 mol of  $\text{CO}_2$  reacts with excess

$\text{MgO}$  what is  $\Delta H$  when 35.5 g of  $\text{CO}_2$  reacts with excess  $\text{MgO}$  this page

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 $2C(s) + O_2(g) \rightarrow 2CO(g)$   $\Delta H = -110.5 \text{ kJ}$  is this endothermic or exothermic?

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 $2H_2(g) + O_2(g) \rightarrow 2H_2O(l)$  energy endothermic the transfer of heat from the surroundings to the system energy  
 $2Hg(s) \rightarrow 2Hg(l) + O_2(g)$

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web  $q = 140 \text{ g} \times 4.184 \text{ J/g} \times 16.0 \text{ }^\circ\text{C}$   $q = 9372 \text{ J}$   $35.5 \text{ g} \times 1 \text{ mole KCN} \times 0.5453 \text{ moles}$   $65.1 \text{ g} \times 0.5453 \text{ mole moles}$  5 when a 15 g sample of  $SnCl_2$  dissolves in 80 g of water in a calorimeter the temperature drops from  $42^\circ\text{C}$  to  $15^\circ\text{C}$  calculate the  $\Delta H$  for the process  
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