

1. Draw the Lewis Structure for NO



2. Draw the Lewis Structure for NOF



3. What is the structural geometry of the NOF molecule
bent

4. The rate constant (k) for a reaction was measured as a function of a temperature. A plot of $\ln k$ versus $1/T$ (in K) is linear and has a slope of -7002 K . Calculate the activation energy for the reaction

$$k = e^{-E_a/R} (+)$$

↑ slope

$$\begin{aligned} -7002 &= -E_a/R \\ -7002 \cdot 8.314 &= -E_a \\ -58210 &= -E_a \\ E_a &= 5.821 \times 10^4 \text{ J/mol} \end{aligned}$$

5. The gas phase rxn of NO w/ F_2 to form NOF and F has an E_a of 6.3 kJ/mol and a frequency factor of $A = 6.00 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$. The rxn is believed to be bimolecular:

$\text{NO}(g) + \text{F}_2(g) \rightarrow \text{NOF}(g) + \text{F}(g)$
What NO is a free radical. Free radicals, species with an odd number of unbound electrons, tends to be more reactive. This is one reason why NO reacts so readily with gases such as fluorine or chlorine.

What is the rate constant at 665°C ?

$$\begin{aligned} 665 + 273 &= 938 \text{ K} \\ k &= 6.00 \times 10^8 \cdot e^{(-6300 / 8.314 \cdot 938)} \\ k &= 2.67 \times 10^8 \end{aligned}$$

6. The rate constant of a chemical reaction increased from 0.100 s^{-1} to 2.80 s^{-1} upon raising the temperature from 25.0°C to 53.0°C .

a. Calculate the value of $(1/T_1 - 1/T_2)$ where T_1 is the initial and T_2 is the final.

$$\begin{aligned} 25 + 273 &= 298 \text{ K} \\ 53 + 273 &= 326 \text{ K} \\ 1/298 - 1/326 &= 2.88 \times 10^{-4} \text{ K}^{-1} \end{aligned}$$

Mastering Chemistry Problems

TD Snyder



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