

INTRODUCTION

Chemistry is the study of matter & change

Mixtures - combination of two or more substances

Scientific notation

Scientific notation is used to express very large or very small numbers

DIMENSIONAL ANALYSIS

BASE

Substance that reacts with acids to form water and a salt



Acid

Substance that reacts with bases to form water and a salt

thermochemistry

Study of heat change in chemical reactions

ENERGY, WORK & HEAT

OXIDATION-REDUCTION

DALTON'S LAW OF PARTIAL PRESSURES

kinetic theory GASES

General Chemistry 1 Notes

30 PAGES

BALANCING CHEMICAL REACTIONS

Change coefficients (numbers in front of formulas) to make number of atoms of each element equal on both sides

PERCENT YIELD

Tells chemist how successful the reaction was

$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

Acid-base reaction

Acids normally start with H or HO

ACIDS vs BASES

ACIDS	WHETHER HAVE
acidic taste	hydrogen
turn litmus red	
conduct electricity	
react with bases	

MOLECULAR EQUATIONS, IONIC EQUATIONS, NET IONIC EQUATIONS

Molecular Equations

Formulas written as if an ionic species existed as molecules



steps for writing ionic & net ionic equations

1. Write balanced molecular equation for the reaction, correct formulas & solubility

HESS'S LAW: When reactants are converted into product, change in enthalpy is the same whether the reaction occurs in one or multiple steps

$$\Delta H_{\text{rxn}} = \sum n \Delta H_f^{\circ}(\text{products}) - \sum m \Delta H_f^{\circ}(\text{reactants})$$

↑ substitution coefficients from balanced reaction

Enthalpy of formation



= zero



SPECIFIC HEAT CAPACITY

Heat capacity of a body is the amount of heat required to raise its temperature by one unit

Notes On General Chemistry

JL Elias



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