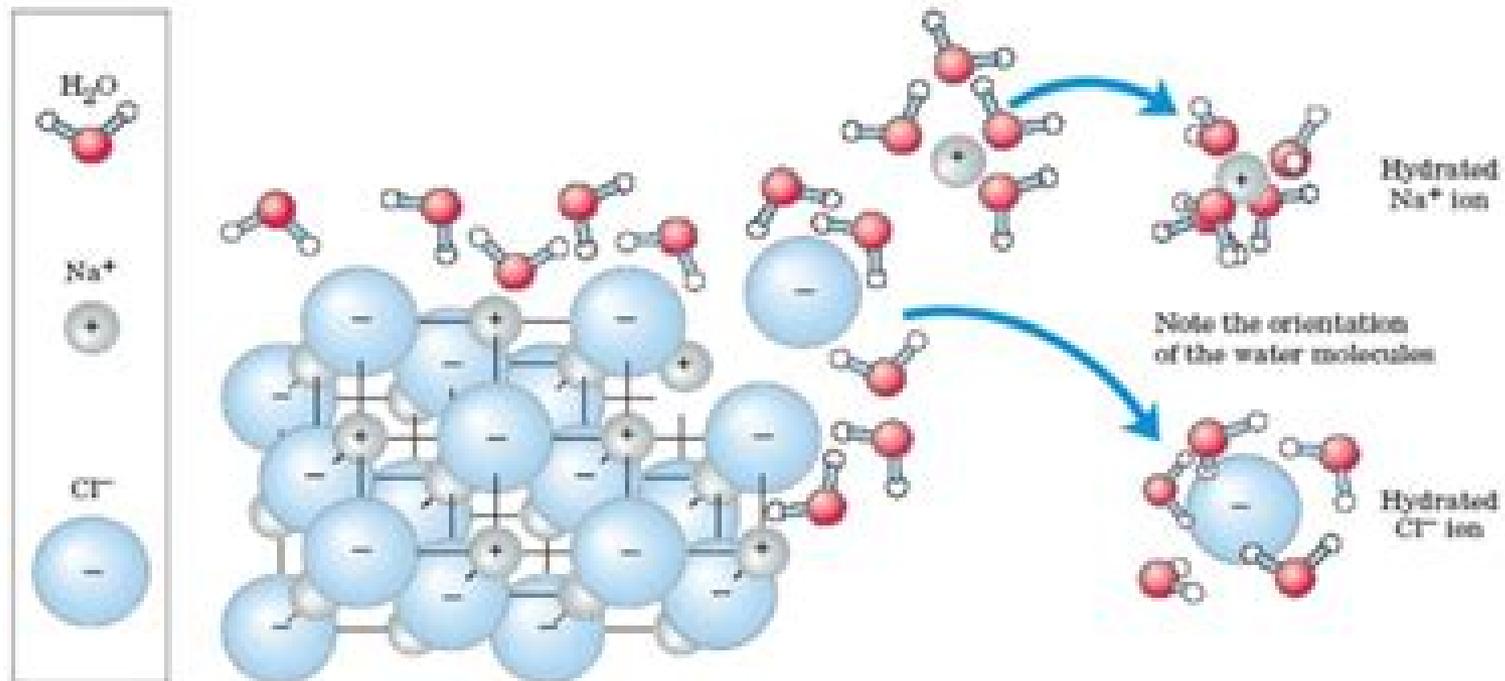


# Entropy Increases as Crystalline Substances Dissolve



**Water as solvent.** Water dissolves many crystalline salts by hydrating their component ions. The NaCl crystal lattice is disrupted as water molecules cluster about the Cl and Na ions. The ionic charges are partially neutralized, and the electrostatic attractions necessary for lattice formation are weakened.

$\Delta G = \Delta H - T\Delta S$ , where  $\Delta H$  has a small positive value and  $T\Delta S$  a large positive value; thus  $\Delta G$  is negative.

# Polar Molecules In Crystalline And Surface Environments From First Principles

**Sebastian Brünink**



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