

Phase Equilibria in Multicomponent Water–Salt Systems

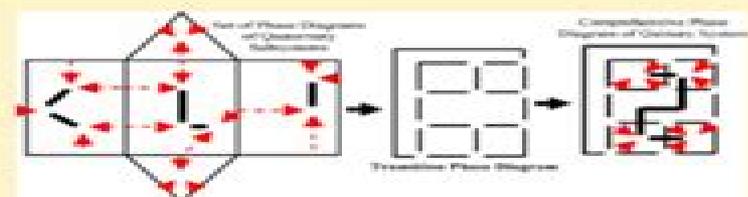
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ABSTRACT: The translation method of determining phase equilibria and construction of phase diagrams for multicomponent systems was applied to investigate the phase equilibria in two quaternary Na, NH₄//SO₄, Cl–H₂O (at 100 °C, 10 °C) and Na₂SO₄–K₂SO₄–MgSO₄–H₂O (at 35 °C) and three quinary Li, K//Cl, CO₃, B₄O₅–H₂O (at 0 °C), Na, K//Cl, SO₄, B₄O₅–H₂O (at 50 °C), and Na, K//Cl, CO₃, B₄O₅–H₂O (at 25 °C) water–salt systems. The comprehensive closed phase equilibria diagrams for each of the systems were constructed at the relevant temperatures. The constructed diagrams for the quinary systems were fragmented into trivariant crystallization volumes for each of the available individual solid phases in the systems. The available experimental results on each of the three investigated quinary systems compose only a portion of the phase equilibria data obtained by us using the techniques of translation method. Comparison of our results with the available literature data for the two quinary Li, K//Cl, CO₃, B₄O₅–H₂O (at 0 °C) and Na, K//Cl, SO₄, B₄O₅–H₂O (at 50 °C) systems shows agreement.



1. INTRODUCTION

While the multicomponent systems are studied by eliminating existing phases in systems, the vagueness about the rest of the system remains without elucidation. In study of phase equilibria in Na⁺, NH₄⁺//SO₄²⁻, HCO₃⁻, Cl⁻–H₂O; Na⁺, Mg²⁺, K⁺//SO₄²⁻, B₄O₅²⁻–H₂O; and Li⁺, Na⁺, K⁺//CO₃²⁻, B₄O₅²⁻–H₂O systems, the NaHCO₃, MgB₄O₇·9H₂O, and Li₂CO₃ phases were eliminated, respectively.^{1–3} However, the phase equilibria beyond the eliminated solid phases have been neither observed nor properly investigated. If the case is like this even in quinary systems, more complex systems are beyond the capabilities of most of the phase equilibria investigations.

The prominent Russian scientist Nikolai Semenovitch Kurnakov⁴ has significantly contributed to the development of theoretical and experimental investigations of multicomponent systems. The two basic “consistency” and “continuity” principles,⁵ as well as the subject and tasks of the physico-chemical analysis, were formulated by him. As had been observed by Kurnakov, the geometrical figures of the subsystems extend into the composition of the global system by the addition of a new component.⁶ Later, the third compatibility principle was proposed by Goroshchenko⁷ as a theoretical justification of Kurnakov’s latter observation. Every geometrical figure of diagrams of subsystems and the overall system is involved in the phase diagram of the global composition according to the compatibility principle.

The “translation method” which is comprehensively described in the literature,^{8–12} was derived from the latter third principle of the physico-chemical analysis and functions in accordance with Gibbs’ phase rule. The method which is employed for the determination of phase equilibria and construction of closed equilibria phase diagrams for multicomponent systems substantially alleviates the hardships in phase equilibria studies of multicomponent systems.

2. DETERMINATION OF PHASE EQUILIBRIA IN MULTICOMPONENT WATER–SALT SYSTEMS USING TRANSLATION METHOD

The method involves three main “through, unilateral and intermediate”^{13–15} techniques which are employed in this section in the investigation of three different quaternary cases. The selected systems are among the experimentally studied ones,^{16–18} hence enabling comparison of results obtained by the translation method. Each case was chosen to show a different technique of the method.

2.1. Phase Equilibria in the Quaternary Na, NH₄//SO₄, Cl–H₂O System at 100 °C. To predict the phase equilibria and construct a “schematic phase equilibria diagram”^{19–21} for the Na, NH₄//SO₄, Cl–H₂O system at 100 °C only “through

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Phase Equilibria In Multicomponent System

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